

## Molecular Shapes (VSEPR)

Molecule	# Atoms bonded to central atom	# Lone pairs (not bonded)	Molecular shape
CO <sub>2</sub>	2	0	Linear
BF <sub>3</sub>	3	0	Trigonal planar
CH <sub>4</sub>	4	0	Tetrahedral
H <sub>2</sub> O	2	2	Bent
O <sub>3</sub>	2	1	Bent
NH <sub>3</sub>	3	1	Pyramidal

In the space below, draw each of the six molecules listed above.