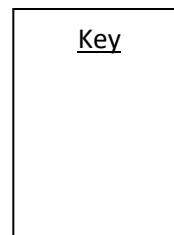
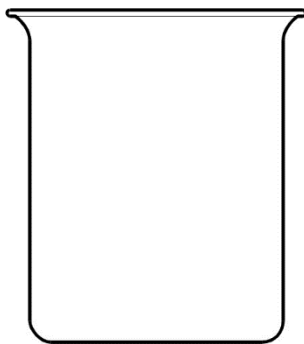
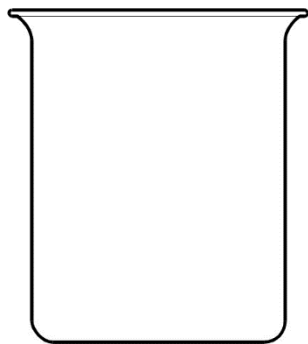


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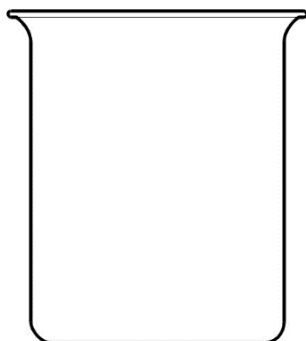
Date: _____

Net Ionic Equations Activity

1. Observe the reaction between aqueous cobalt (II) nitrate and aqueous sodium carbonate. Imagine how they look at the molecular level.
 - a. Write the molecular equation for this reaction:
 - b. Write the balanced net ionic equation for this reaction:
 - c. What is/are the spectators in this reaction?
 - d. What is the precipitate's name?
 - e. Draw the reactants before the reaction. (You need at least 3 molecules/ions of each substance present in each beaker and be mindful of the placement of your drawings.)



- f. Draw the products after the reaction. (Be aware of how many of each ion you started with.)



2. Observe the reaction between aqueous Barium chloride and sodium sulfate. Imagine how they look at the molecular level.

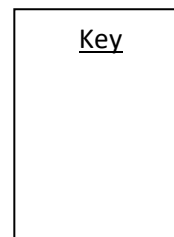
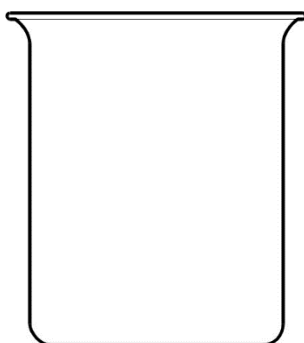
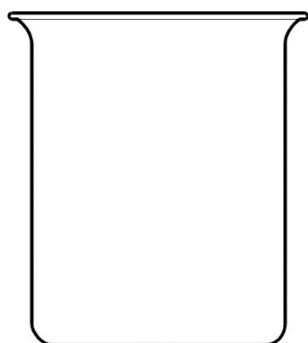
a. Write the molecular equation for this reaction:

b. Write the balanced net ionic equation for this reaction:

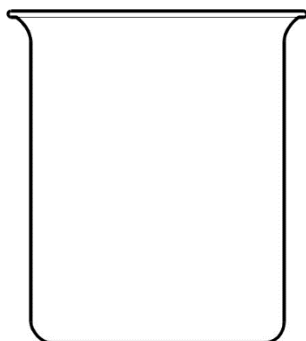
c. What is/are the spectators in this reaction?

d. What is the precipitate's name?

e. Draw the reactants before the reaction. (You need at least 3 molecules/ions of each substance present in each beaker and be mindful of the placement of your drawings.)



f. Draw the products after the reaction. (Be aware of how many of each ion you started with.)



3. Observe the reaction between aqueous copper (II) chloride and sodium carbonate. Imagine how they look at the molecular level.

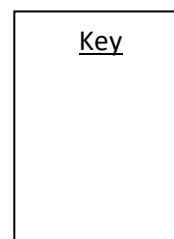
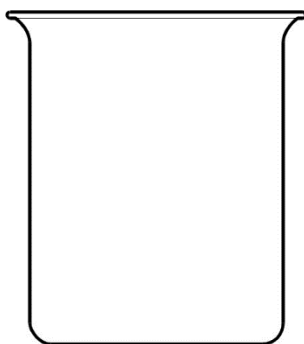
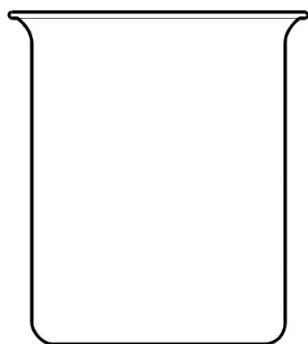
a. Write the molecular equation for this reaction:

b. Write the balanced net ionic equation for this reaction:

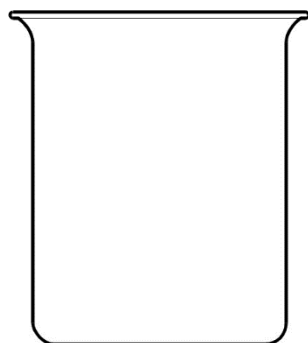
c. What is/are the spectators in this reaction?

d. What is the precipitate's name?

e. Draw the reactants before the reaction. (You need at least 3 molecules/ions of each substance present in each beaker and be mindful of the placement of your drawings.)



f. Draw the products after the reaction. (Be aware of how many of each ion you started with.)



4. Observe the reaction between aqueous lead (II) nitrate and sodium iodide. Imagine how they look at the molecular level.

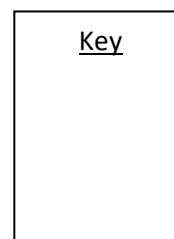
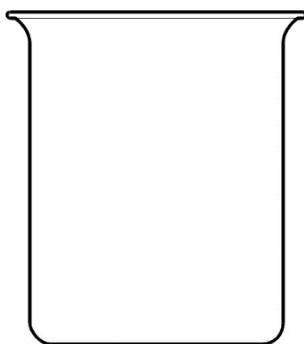
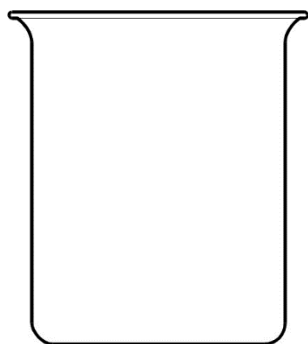
a. Write the molecular equation for this reaction:

b. Write the balanced net ionic equation for this reaction:

c. What is/are the spectators in this reaction?

d. What is the precipitate's name?

e. Draw the reactants before the reaction. (You need at least 3 molecules/ions of each substance present in each beaker and be mindful of the placement of your drawings.)



f. Draw the products after the reaction. (Be aware of how many of each ion you started with.)

