

## Nomenclature Practice

### Part I

**Identify** each compound as either ionic or molecular. Then **name** it and determine the molar mass.

<u>Compound</u>	<u>Ionic/Molecular</u>	<u>Name</u>	<u>Molar Mass</u>
1. MgO			
2. K <sub>3</sub> P			
3. SO <sub>2</sub>			
4. Fe <sub>2</sub> S <sub>3</sub>			
5. N <sub>2</sub> O <sub>5</sub>			
6. BaSO <sub>4</sub>			
7. Zn(NO <sub>3</sub> ) <sub>2</sub>			
8. CCl <sub>4</sub>			
9. SF <sub>6</sub>			
10. Ni <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>			

### Part II

**Identify** each compound as either ionic or molecular. Then **write** a correct formula and determine the molar mass.

<u>Compound</u>	<u>Ionic/Molecular</u>	<u>Name</u>	<u>Molar Mass</u>
11. magnesium sulfate			
12. nitrogen triiodide			
13. lead (II) phosphate			
14. ammonium carbonate			
15. dichlorine monoxide			
16. carbon dioxide			
17. aluminum chloride			
18. chromium (III) oxide			
19. potassium iodide			
20. diphosphorus tetraoxide			