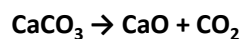


Percent Yield Notes and Practice

Percent Yield

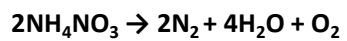
- % Yield = $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$
- Actual—amount actually formed and measured in lab
- Theoretical—amount calculated using stoichiometry calculation

(2)



- What is the theoretical yield of calcium oxide if 24.8 g calcium carbonate is heated?
- What is the percent yield if Billy the chemistry student produces 13.1 g CaO in lab?

(3)



- Suzy the chemist performs this reaction. She starts with 228 grams of ammonium nitrate and ends up with 26.8 liters of nitrogen (assume STP). What is her percent yield?

4
