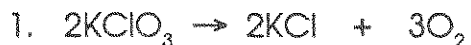
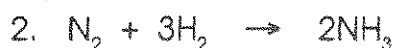


STOICHIOMETRY: MASS-MASS PROBLEMS

Name _____



How many grams of potassium chloride are produced if 25 g of potassium chlorate decompose?



How many grams of hydrogen are necessary to react completely with 50.0 g of nitrogen in the above reaction?

3. How many grams of ammonia are produced in the reaction in Problem 2?

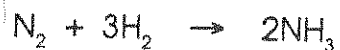


How many grams of silver chloride are produced from 5.0 g of silver nitrate reacting with an excess of barium chloride?

5. How much barium chloride is necessary to react with the silver nitrate in Problem 4?

STOICHIOMETRY: VOLUME-VOLUME PROBLEMS

Name _____



What volume of hydrogen is necessary to react with five liters of nitrogen to produce ammonia? (Assume constant temperature and pressure.)

2. What volume of ammonia is produced in the reaction in Problem 1?



If 20 liters of oxygen are consumed in the above reaction, how many liters of carbon dioxide are produced?



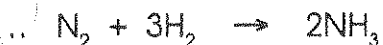
If 30 mL of hydrogen are produced in the above reaction, how many milliliters of oxygen are produced?



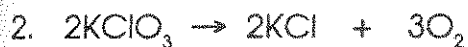
How many liters of carbon dioxide are produced if 75 liters of carbon monoxide are burned in oxygen? How many liters of oxygen are necessary?

STOICHIOMETRY: MIXED PROBLEMS

Name _____



What volume of NH_3 at STP is produced if 25.0 g of N_2 is reacted with an excess of H_2 ?



If 5.0 g of KClO_3 is decomposed, what volume of O_2 is produced at STP?

3. How many grams of KCl are produced in Problem 2?



What volume of hydrogen at STP is produced when 2.5 g of zinc react with an excess of hydrochloric acid?



How many molecules of water are produced if 2.0 g of sodium sulfate are produced in the above reaction?



If 10.0 g of aluminum chloride are decomposed, how many molecules of Cl_2 are produced?
